

2,3-Dichloro-5,6-dicyano-1,4-benzoquinone-Catalyzed Reactions Employing MnO_2 as a Stoichiometric Oxidant

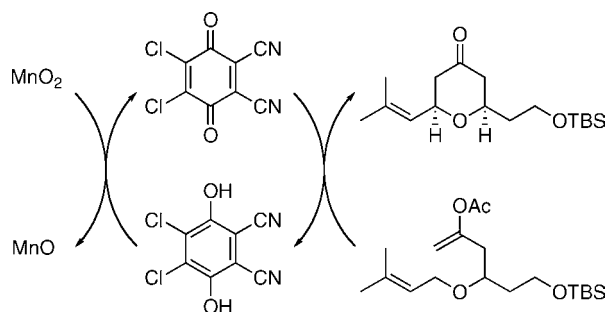
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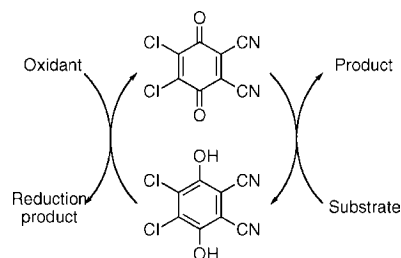
ABSTRACT



Several oxidative reactions can be effected with MnO_2 in the presence of substoichiometric quantities of DDQ. These transformations include oxidative cyclization, deprotection, and dehydrogenation reactions. The use of MnO_2 as a terminal oxidant for DDQ-mediated reactions is attractive based on economical and environmental factors.

DDQ (2,3-dichloro-5,6-dicyano-1,4-benzoquinone) is a highly effective oxidant for a number of chemical transformations including protecting group removal, aromatization, acetal formation, and biaryl construction.¹ DDQ is moderately expensive, however, with a cost of \$526/mol according to the 2009–2010 Aldrich catalog.² In addition to concerns about cost, DDQ also poses modest toxicity concerns, with an LD_{50} of 82 mg/kg,³ and the potential for HCN liberation upon exposure to H_2O . These issues, coupled with the potential purification difficulties associated with utilizing stoichiometric amounts of an organic oxidant, have led to the search for economical and environmentally benign oxidants that regenerate DDQ from its reduced hydroquinone form (Scheme 1). HNO_3 is a suitable oxidant for this

Scheme 1. DDQ as a Catalytic Oxidant



purpose,⁴ though its strong acidity is often not compatible with functionalized organic compounds and necessitates that the reoxidation be conducted as an independent step. Chandrasekhar and co-workers reported⁵ that DDQ could

(1) Buckle, D. R. In *Encyclopedia of Reagents for Organic Synthesis*; Paquette, L. A., Ed.; John Wiley & Sons: Chichester, UK, 1995; Vol. 3, p 1699.

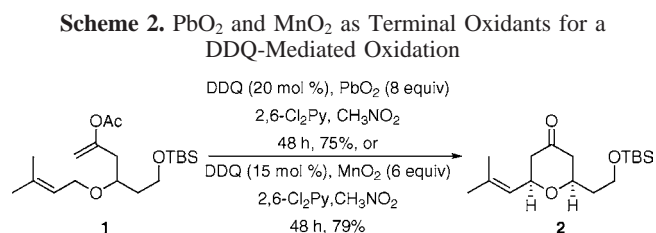
(2) All price quotes are based on 100 g bottles of reagent grade material.

(3) Rat LD_{50} values can be found on MSDS sheets.

(4) Newman, M. S.; Khanna, V. K. *Org. Prep. Proced. Int.* **1985**, *17*, 422.

be used as a catalytic oxidant with FeCl₃ serving as the stoichiometric oxidant. This system is economically and ecologically attractive, with FeCl₃ costing only \$41/mol, and having an LD₅₀ of 450 mg/kg, though the Lewis acidity of FeCl₃ again makes it incompatible with many functional groups. Sharma and co-workers showed⁶ that Mn(OAc)₃ is also an effective agent for DDQ regeneration that does not exhibit strong Lewis acidity. However, Mn(OAc)₃·2H₂O is actually more expensive than DDQ at a price of \$647/mol. The cost factor is compounded by the fact that Mn(OAc)₃ is a single electron oxidant, making 2 equiv necessary to regenerate DDQ.⁷

In accord with our efforts⁸ and those of other groups⁹ to use DDQ as a reagent for oxidative carbon–carbon bond forming reactions and our interest in developing catalytic oxidative transformations,¹⁰ we have initiated an effort to identify an inexpensive, nonacidic, and environmentally benign reagent that can be used as a terminal oxidant in the presence of substoichiometric DDQ loadings. Following the observation that the use of FeCl₃ as a terminal oxidant for DDQ-catalyzed carbon–hydrogen bond functionalization reactions induced substrate decomposition, we directed our initial screen for terminal oxidants toward metal oxides that are known to effect phenol oxidations. This search led us to explore PbO₂¹¹ (\$117/mol, LD₅₀ = 220 mg/kg). Exposing prenyl ether **1** to DDQ (20 mol %) and PbO₂ in CH₃NO₂ provided tetrahydropyrone **2** in 75% yield after 48 h (Scheme 2). Incorporating 2,6-dichloropyridine made this reaction and



subsequent transformations proceed more efficiently because it can act as a nonoxidizable base and because it quenches

(5) Chandrasekhar, S.; Sumithra, G.; Yadav, J. S. *Tetrahedron Lett.* **1996**, 37, 1645.

(6) Sharma, G. V. M.; Lavanya, B.; Mahalingam, A. K.; Krisha, P. R. *Tetrahedron Lett.* **2000**, 41, 10323.

(7) For a review of Mn(OAc)₃ chemistry, see: Snider, B. B. *Chem. Rev.* **1996**, 96, 339.

(8) (a) Tu, W.; Liu, L.; Floreancig, P. E. *Angew. Chem., Int. Ed.* **2008**, 47, 4184. (b) Tu, W.; Floreancig, P. E. *Angew. Chem., Int. Ed.* **2009**, 48, 4567. (c) Liu, L.; Floreancig, P. E. *Org. Lett.* **2009**, 11, 3152. (d) Jung, H. H.; Floreancig, P. E. *Tetrahedron* **2009**, 65, 10830. (e) Liu, L.; Floreancig, P. E. *Angew. Chem., Int. Ed.* **2010**, 49, 3069. (f) Cui, Y.; Tu, W.; Floreancig, P. E. *Tetrahedron* **2010**, 66, 4867. (g) Liu, L.; Floreancig, P. E. *Angew. Chem., Int. Ed.* **2010**, 49, 5894.

(9) (a) Ying, B.-P.; Trogden, B. G.; Kohlman, D. T.; Liang, S. X.; Xu, Y.-C. *Org. Lett.* **2004**, 6, 1523. (b) Zhang, Y.; Li, C.-J. *Angew. Chem., Int. Ed.* **2006**, 45, 1949. (c) Li, Z.; Li, H.; Guo, X.; Cao, L.; Yu, R.; Li, H.; Pan, S. *Org. Lett.* **2008**, 10, 803. (d) Cheng, D.; Bao, W. *J. Org. Chem.* **2008**, 73, 6881. (e) Tsang, A. S.-K.; Todd, M. H. *Tetrahedron Lett.* **2009**, 50, 1199. (f) Yu, B.; Jiang, T.; Li, J.; Su, Y.; Pan, X.; She, X. *Org. Lett.* **2009**, 11, 3442. (g) Benfatti, F.; Capdevila, M. G.; Zoli, L.; Benedetto, E.; Cozzi, P. G. *Chem. Commun.* **2009**, 5919. (h) He, Z.; Lin, X.; Zhu, Y.; Wang, Y. *Heterocycles* **2010**, 81, 965. (i) Guo, C.; Song, J.; Luo, S.-W.; Gong, L.-Z. *Angew. Chem., Int. Ed.* **2010**, 49, 5558.

the acylium ion that forms when the enol acetate group reacts with the oxocarbenium ion. Cognizant of the long-term toxicity of lead, we explored other metal oxides as terminal oxidants. Exposing **1** to 15 mol % DDQ and MnO₂ in CH₃NO₂ provided **2** in 79% yield. While the reaction time was long (48 h) this result was quite gratifying in consideration of the low cost of MnO₂ (\$28/mol) and its negligible toxicity (LD₅₀ = 3478 mg/kg). The products of these reactions were quite easy to isolate, with the excess metal-containing reagents and byproduct being removed with a simple filtration over silica gel. In general oxidative cyclizations are slower in CH₃NO₂ than in less polar solvents, such as 1,2-dichloroethane. However, the regeneration of DDQ was clearly faster in CH₃NO₂, making it the solvent of choice for this process. No conversion of **1** to **2** was observed when PbO₂ or MnO₂¹² was used in the absence of DDQ, confirming that the reactions were mediated by DDQ and that the metal oxides function to oxidize the hydroquinone. Additional evidence in support of DDQ regeneration was provided by observing the formation of DDQ upon subjecting 2,3-dichloro-5,6-dicyanohydroquinone¹³ to MnO₂ in CH₃NO₂.¹⁴

The successful demonstration of MnO₂-mediated DDQ turnover for our oxidative cyclization protocol led us to examine the generality of the protocol. A representative group of these reactions is shown in Table 1. Allylic ethers, benzylic ethers, and vinyl oxazolidinones serve as substrates for the protocol. While we did not study a broad range of nucleophiles, we observed that alcohols are tolerated by the reaction conditions in addition to electron-rich alkenes. Quaternary centers can be formed in the process, and sterically hindered ethers can be used as substrates. In general the yields of the reactions were within approximately 10% of the yields for the corresponding reactions that employ 2.0 equiv of DDQ. The rates of these reactions roughly correlate to the corresponding stoichiometric reactions, though the long reaction times that are observed for even the most reactive substrates show that hydroquinone oxidation is the rate-limiting step. The only significant problem was encountered when slowly reacting substrate **19** (entry 9) was examined. This problem can be attributed to the change in solvent, because these cyclizations proceed more slowly in CH₃NO₂ than DCE, and to the low reaction rates that result from employing a substoichiometric amount of oxidant.

We examined the applicability of using MnO₂ as a terminal oxidant to other commonly employed DDQ-mediated transformations to explore the generality of the process. A number of these reactions are shown in Scheme 3. PMB ether cleavage can be effected, as seen in the conversion of **20** to **21**. Water deactivates MnO₂, so MeOH was used as the nucleophile in this transformation. Oxidizing hydroxyl-

(10) Kumar, V. S.; Aubele, D. L.; Floreancig, P. E. *Org. Lett.* **2001**, 3, 4123.

(11) Cook, C. D.; English, E. S.; Wilson, B. J. *J. Org. Chem.* **1958**, 23, 755.

(12) For an example of a manganese-catalyzed ether oxidation, see: Kamijo, S.; Amaoka, Y.; Inoue, M. *Chem. Asian J.* **2010**, 5, 486.

(13) van Haeringen, C. J.; West, J. S.; Davis, F. J.; Gilbert, A.; Hadley, P.; Pearson, S.; Wheldon, A. E.; Henbest, R. G. C. *Photochem. Photobiol.* **1998**, 67, 407.

(14) We thank an anonymous reviewer for suggesting this experiment.

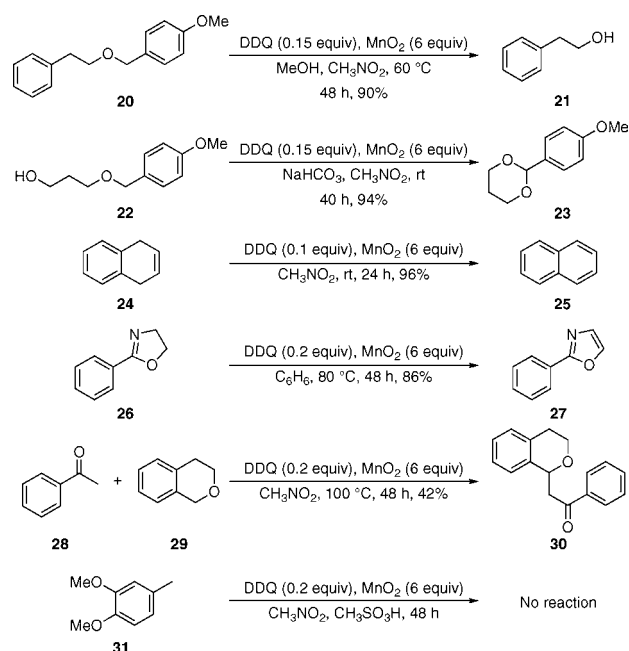
Table 1. DDQ-Catalyzed Oxidative Cyclization Reactions^a

entry	substrate	product	time (h)	yield (%) ^b
1			41	76
2			44	83
3			40	92
4 ^c			48	75
5			36	68
6			24	87
7			40	70
8			30	81
9		-	-	-

^a Representative procedure: DDQ was added in three equal portions over the course of the reaction (0.15 equiv total) to a suspension of the substrate, 2,6-dichloropyridine (2 equiv), and activated MnO₂ (6 equiv) in CH₃NO₂ (0.1 M). The reaction was stirred at room temperature for the indicated period of time. ^b Yields refer to isolated, purified products. ^c Reaction required 0.2 equiv of DDQ.

containing PMB ether **22** in the absence of MeOH provided PMP acetal **23**. Dehydrogenation reactions are quite effective, with dihydronaphthalene (**24**) being converted to naphthalene (**25**) in 96% yield. This process can also be applied to the conversion of oxazolines to oxazoles, as seen in the conversion of **26** to **27**.¹⁵ This method of oxazole preparation is an improvement over a previously reported variant¹⁶ in which specially prepared high surface area MnO₂ (25 equiv) is used for oxazoline oxidation in the absence of DDQ,¹⁷ and should

(15) For recently reported oxidative approaches to oxazole synthesis, see: (a) Graham, T. H. *Org. Lett.* **2010**, *12*, 3614. (b) Phillips, A. J.; Uto, Y.; Wipf, P.; Reno, M. J.; Williams, D. R. *Org. Lett.* **2000**, *2*, 1165.

Scheme 3. MnO₂-Mediated DDQ-Catalyzed Reactions

be quite beneficial for economical syntheses of these important structures.¹⁸ The protocol can be utilized for Li's cross dehydrogenative coupling¹⁹ between acetophenone (**28**) and isochroman **29** to form alkylation product **30**. While the yield for the reaction was modest, the transformation illustrated the capability of forming carbon–carbon bonds from readily available precursors by using an inexpensive oxidant. Attempts to conduct a biaryl formation of **31** through Rathore's variant²⁰ of the Scholl reaction were unsuccessful, most likely indicating that the MeSO₃H that is necessary for the coupling deactivates MnO₂.

The ability to conduct large-scale transformations with minimal DDQ loading would be a major benefit of this protocol. We confirmed the capacity to apply the procedure to large-scale reactions by converting **24** to **25** on a 4.1 g (31.5 mmol) scale. The reaction proceeded in comparable yield (94%) to the smaller scale reaction (96%) and reached completion more rapidly (16 h compared to 24 h). The rate enhancement can be attributed to the reaction being conducted at approximately 3 M as opposed to the 0.1 M concentration that was used in the smaller scale transformations.

We have shown that a number of oxidations can be effected by using catalytic DDQ with MnO₂ serving as an inexpensive and environmentally benign terminal oxidant. These transformations include oxidative cyclizations, ether

(16) Aoyama, T.; Sonoda, N.; Yamauchi, M.; Toriyama, K.; Anzai, M.; Ando, A.; Shioiri, T. *Synlett* **1998**, 35.

(17) We did not observe oxazoline oxidation using activated reagent grade MnO₂ in the absence of DDQ.

(18) Jin, Z. *Nat. Prod. Rep.* **2006**, *23*, 464.

(19) Zhang, Y.; Li, C.-J. *J. Am. Chem. Soc.* **2006**, *128*, 4242. For a review of cross dehydrogenative coupling reactions, see: Li, C.-J. *Acc. Chem. Res.* **2009**, *42*, 335.

(20) Zhai, L.; Shukla, R.; Rathore, R. *Org. Lett.* **2009**, *11*, 3474.

cleavages, dehydrogenations, and cross-dehydrogenative couplings. The processes, while significantly slower than the corresponding stoichiometric reactions, are comparable with respect to yield and the resulting products are easier to purify. The protocol is incompatible with very slow oxidations and with the presence of water or strong acid. The range of successful transformations reported herein and the capacity to conduct reactions on a large scale, however, suggest that this process should be useful for most DDQ-mediated oxidations, and the cost advantages and minimal toxicity benefits could prove to be very valuable for multigram syntheses.

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Supporting Information Available: Experimental details and characterization of all new compounds. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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